

Niobium doped BGO glasses: Physical, Thermal and Optical Properties

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Abstract: IR-transparent niobium substituted heavy metal oxide glass system in formula composition (100-x) BGO75- xNb₂O₅; 5 ≤ x ≤ 25 was fabricated by quenching and press molding technique. Glassy state was confirmed by XRD. Density ρ_{exp} varied 6.381 g/cc-7.028 g/cc ±0.06%. Modifying behavior of Nb₂O₅ was corroborated by rate of increase in theoretical volume V_{th} , measured volume V_{exp} and oxygen molar volume V_{MO}^{-2} . Nb⁵⁺ had larger cation radius and greater polarizing strength as compared to Ge⁴⁺ ions. It replaced Ge⁴⁺ sites introducing more NBOs in the network. Transformation temperatures T_g , T_x and T_{p1} were 456-469 °C ±2 °C, 516-537 °C ±2 °C and 589-624 °C ±2 °C respectively. In the range from room temperature to 400 °C the coefficient of linear thermal expansion α was $5.431 \pm 0.001 * 10^{-6} K^{-1}$ to $7.333 \pm 0.001 * 10^{-6} K^{-1}$. $\Delta T = T_x - T_g$ and $\Delta T_{p1} = T_{p1} - T_g$ varied collinearly with increase in niobium concentration and revealed thermal stability against devitrification. The direct bandgap E_g values lay in 3.24-2.63 eV ±0.01 eV range and decreased due to impurity states of Nb⁵⁺ within the forbidden band. Mobility edges obeyed Urbach law verifying amorphousness of the compositions. The IR-transparency window extended from 0.38-6.1 μm in MIR which is wider than binary BGO glass as well as BGO single crystal.

Keywords: BGO75; band gap; IR-transparent; Thermal stability; Urbach law

I. Introduction

Glass is an optically isotropic medium. It is an amorphous material with transparency window engineered via flexibility of composition. The unique class is oxide glasses. Their chemo-physical characteristics develop due to varying covalent, ionic or mixed covalent-ionic bonding nature of oxygen net within the cationic environment set up by the ratio of the cations [1]. There exists a great interest in ceramics fabricated via glass route. These transparent or translucent ceramics are clusters of more or less ordered nano/micro motifs embedded in amorphous matrix [2]. Transparent glasses on ceramization at the expense of some transparency may inject and multitude unique photonic, electro-optic, piezo-optic and/or magneto-optic properties.

Bi₃Ge₄O₁₂ (BGO) are synthetic eulytine crystals. Fabrication of BGO single crystal is possible only with highly pure oxides and exact stoichiometric control of composition and thermal history [3]. It is of great importance in photonic materials and has diverse technological applications [2-6]. The binary system Bi₂O₃-GeO₂ exhibits glassy state in limited range of composition in GeO₂ rich part of composition [4]. BGO glass ceramics are relatively easy to fabricate and cost effective [7]. These ceramics due to their technologically important features like IR-transmission, luminosity around 1.5μm and non-linear properties can replace BGO single crystals [8]. It is of great technical interest to probe the Bi₂O₃-GeO₂ system with third oxide like Nb₂O₅ which could enhance its glass formation boundaries, foster its optical response in MIR and may provide control over ceramization of BGO crystallites [9].

The objective of the present study is to report the glass formation in (100-x) [2.5Bi₂O₃-7.5GeO₂ (BGO75)] - x Nb₂O₅; 5 ≤ x ≤ 25 system with Nb₂O₅ replacement and lit upon the physical and opto-thermal properties of the ternary system. In this system GeO₂ is a classical network former, Bi₂O₃ a conditional glass former and Nb₂O₅ is incorporated as transition metal dopant that could bring about thermo-mechanical stability and network reinforcement extending the transmission window [7]. Such a study is of academic as well as technical interest as BGO eulytine structure single crystal and binary BGO ceramics are being explored extensively due to their opto-electronic applications for WDM and radiation detectors/absorbers. In authors' knowledge ab-initio study on ternary glass system niobium doped BGO75 glasses is reported for the first time.

II. Experimental Procedures

2.1. Sample Fabrication

Glass samples (10g each) of nominal composition (100-x) (BGO75) -xNb₂O₅; 5 ≤ x ≤ 25 in mole % ratio of oxides were prepared by melt-quench technique. Analytical grade oxides (Sigma Aldrich: 99.9% or 99.99% purity; with trace metal basis) were used. Powders were mixed thoroughly in a quartz mortar and pestle for half an hour. Admixture was put in platinum crucible and dried in box-type furnace for one hour at 400°C in order to remove moisture. Then the crucible was shifted to another open air pre-heated furnace at temperature

1323K to 1443K with varying composition of the sample. Fused was homogenized for 20-30 min. Circular sample disks were casted using a preheated stainless steel mold which was pre-annealed for one hour at 200 °C to avoid thermal shock due to large temperature difference of melt and mold. Fabricated samples were annealed at 350 °C for one hour and furnace cooled to room temperature. Glasses varied in yellow shade to brownish tint as the niobium concentration and soaking temperature/time was increased. Disk samples were wet polished mechanically to optical finish using carbide papers of grit size 400-2000 gradually. Samples withstood mechanical stability while polishing and grinding.

2.2. X-ray Diffraction (XRD)

XRD pattern of the fabricated samples were measured using Geiger flex D/Max-3D M/S Rigaku Corporation, Japan. It exploited nickel filter and copper target. CuK α radiation with $\lambda = 1.542$ nm was used at power rating 35 kV x 30 mA. XRD pattern was scanned in the range 5° to 80° of 2 θ with a step size of 0.05°. Plano-parallel disks were used for XRD measurements.

2.3. Density

Density of the bead samples was measured exploiting Archimedes' buoyant liquid method using Idoethane (density 1.595 g/cc) as immersion liquid with single pan electronic balance of accuracy ± 0.00001 g. Error in density value was less than $\pm 0.06\%$ and results were reproducible to $\pm 0.01\%$. Density was calculated using the relation

$$d_{exp} = \frac{m_1}{(m_1 + m_2) - m_3} * d_{liq} \quad (1)$$

Where m_1 is the mass of the sample in air, m_2 is the mass of the pan in liquid, m_3 is the mass of pan and sample in the liquid and d_{liq} the density of the immersion liquid used. The difference $(m_1 + m_2) - m_3$ gives mass of the equivalent volume of liquid displaced. Density values were confirmed by comparing the values obtained by automated ULTRAPYCNOMETER 1000 from Quantachrome Instruments with result accuracy $\pm 0.03\%$ and repeatability $\pm 0.01\%$. The average volume of each sample immersed in argon purged cell was measured for three preprogrammed runs. The volume standard deviation was less than 0.0035%. The temperature of the argon purged cell was 26 °C ± 1 °C during the measurement.

2.4. Dilatometry

NETZSCH dilatometer 402E was used to record co-efficient of linear thermal expansion (CTE). Sample powder was heated at the nominal rate 5 °C/min within an alumina crucible from room temperature to 400 °C. CTE (α) was noted from the slope of $\frac{dL}{L_0}$ vs. temperature curve through computer interfacing using the relation

$$\alpha = \frac{1}{dT} \left(\frac{dL}{L_0} \right) \quad (2)$$

2.5. Characteristic Temperatures

DTA-curves of the sample in the range 25 – 850 °C were recorded on STA-PT-1600, Linseis, Germany under static environment. Micro powder sample was heated at nominal rate of 10°C/min. in alumina crucible. T_g the glass transition temperature, T_x the onset of crystallization temperature, T_{pi} for i^{th} crystallization peak temperature were noted from DTA curves. The glass transition temperature T_g denotes the temperature range of an endothermic shift of baseline in the DTA curve. The crystallization onset T_x refers to temperature where exothermic inflection just begins. The temperature of exothermic peak inflection T_p corresponds to maximum crystallization rate. Liquidus temperature T_l is the minimum of endothermic peak apex value in the DTA-curve.

2.6. UV-VIS-NIR spectra

Optically finished samples (disks) were used to record UV-VIS-NIR spectra of the samples exploiting spectrophotometer U-4000, Hitachi with resolution = 2 nm in 200-2500 nm range. Absorbance A and percent transmission T % were recorded independently applying baseline correction and computer interface.

2.7. FTIR spectra

FTIR spectra of some annealed parallel disk samples of thickness 1-2 mm were recorded on model Shimadzu 8400s at room temperature. The recorded energy range was 400-4000 cm^{-1} with resolution exploited to be 2 cm^{-1} . An average of 20 runs was recorded.

III. Results and Discussion

3.1. Composition

Glasses in the system (100-x) [2.5Bi₂O₃-7.5GeO₂ (BGO75)] - x Nb₂O₅ with $5 \leq x \leq 25$ in mole % ratio were fabricated. The network former oxide GeO₂ and conditional glass former Bi₂O₃ had fixed proportion

75:25. This ratio lay within the glass forming GeO_2 rich region of the binary Bi_2O_3 - GeO_2 phase diagram [4]. 2d transition metal niobium penta oxide (Nb_2O_5) replaced the binary in gradual steps of 5 mol%. The glasses were named accordingly BGO75-Nbx as reported in **Table1** with the mol % ratio of Nb_2O_5 in each glass.

3.2. XRD

Fig.1 shows x-ray diffraction pattern of five niobium doped BGO samples. All compositions exhibited a broad diffraction envelope. This confirms the amorphous state of the samples in BGO75-Nb glasses.

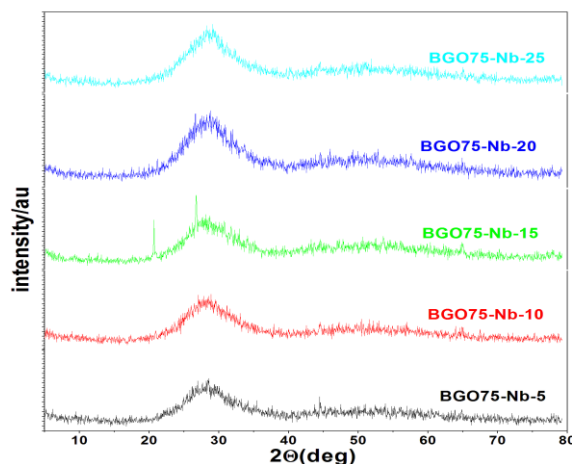


Fig. 1 XRD-pattern of BGO75-Nb glasses

3.3. Density and Molar volume

The simplest way of detecting structural changes in glass network, with compositional variation is by probing its rigidity/flexibility by means of density, molar volume and/or oxygen molar volume of the glass system. Theoretically density is a linear function of composition and can be estimated additively for glass as $\rho_{th} = \frac{\sum x_i \rho_i}{\sum n_i}$; where ρ_i is the density of i^{th} oxide component and x_i is its mole fraction in the composition. Molar volume V_M is calculated by dividing the molar mass of the composition by its density as $V_M = \frac{\sum x_i m_i}{d}$ and oxygen molar volume V_{MO}^{-2} of the network of oxygen in one mole of glass composition can be calculated using the relation $V_{MO}^{-2} = V_M * \frac{1}{\sum x_i n_i}$; where n_i denotes number of oxygen atoms contributed by each oxide [7]. It is volume fraction occupied by anions in the network.

In **Table1** are reported the experimentally measured and theoretically estimated densities ρ_{exp} , ρ_{th} and corresponding molar volume $V_{M exp}$, $V_{M th}$ along with oxygen molar volume V_{MO}^{-2} . Values obtained by using pycnometric measurement ρ_{pyk} and $V_{M pyk}$ are also included in **Table1**.

Table1 Glass code, mol% concentration, densities calculated theoretically ρ_{th} , measured by Archimedes' principle ρ_{exp} , using pycnometer ρ_{pyk} , corresponding molar volumes V_{th} , V_{exp} , V_{pyk} and oxygen molar volume V_{mO}^{-2} of (100-x)BGO75-Nbx glasses

composition code	Nb_2O_5 mol%	d_{Exp} g/cc	V_{exp} cc	d_{th} g/cc	V_{th} cc	d_{pyk} g/cc	V_{pyk} cc	V_{mO}^{-2} cc
BGO75-Nb05	0.52	7.028	27.73	5.39	36.15	8.4278	23.13	15.324
BGO75-Nb10	1.10	6.384	30.53	5.39	36.17	5.9908	32.53	17.826
BGO75-Nb15	1.73	6.469	30.11	5.38	36.2	6.3409	30.72	18.458
BGO75-Nb20	2.44	6.381	30.52	5.38	36.22	6.3766	30.54	19.688
BGO75-Nb25	3.23	6.384	30.50	5.37	36.25	6.0779	32.03	20.765

Fig.2 presents the results in graphical form as a function of increasing niobium-oxide concentration in mol%. Theoretically the density of the glasses decreased from 5.39 g/cc to 5.37 g/cc with increasing mol% concentration of Nb_2O_5 . It was expected as lighter mass cations replaced the relatively heavier average cation in the BGO75 glass. But experimentally this regular monotonic decrease was not observed. The density values exploiting liquid displacement method ρ_{exp} varied 6.381 g/cc -7.028 g/cc $\pm 0.06\%$ whereas density measured using pycnometer was 5.9908 g/cc -8.4278 g/cc $\pm 0.03\%$. It can be noted that the densities measured using two different methods are strongly correlated and comparable in magnitude but higher than those estimated through

formula composition. The difference in two measured values of densities may be due to rounding off the mass values of samples to four significant figures for Pycnometric volume measurements.

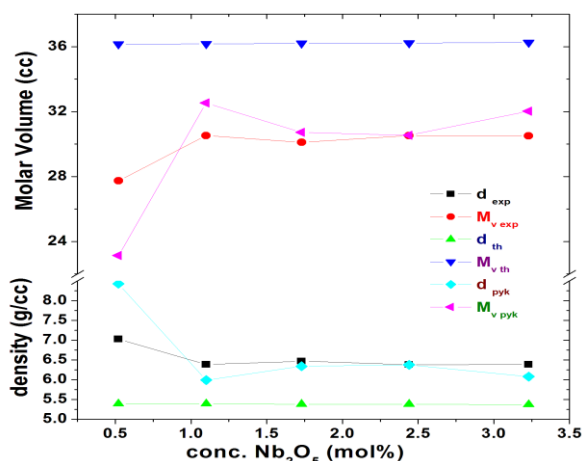


Fig. 2 Variation in density and molar volume with mol% rise in Nb₂O₅ concentration for BGO75 glasses

The density values show that the glass becomes more rigid and compact as niobium is introduced in the composition. This may be due to the fact that ionic radii of cations are in the order Bi³⁺ > Nb⁵⁺ > Ge⁴⁺ [10]. Also the polarizing strength of niobium ion is less than that of bismuth ion but greater than Ge⁴⁺ [10, 11]. These factors may result in change in coordination polyhedra of the network. Nb⁵⁺ ion can occupy some positions within the network as modifier replacing Ge⁴⁺ atom or ion exchange with Bi³⁺ ions. Resulting increase in molar volume occupied is an indication that it replaces Ge sites. Logically, molar volume increases with (i) the no. of oxygen atoms (ii) concentration and radius of cation (iii) rise in co-ordination number (iv) decrease in polarizability of cation [12]. These factors collectively exhibited monotonic increase in volume. Molar volume and Oxygen molar volume both increased with increasing mol% concentration of Nb₂O₅ initially up to Nb10. This implied modifying behavior of Nb⁵⁺ ions incorporating NBOs and depolymerization in the system.

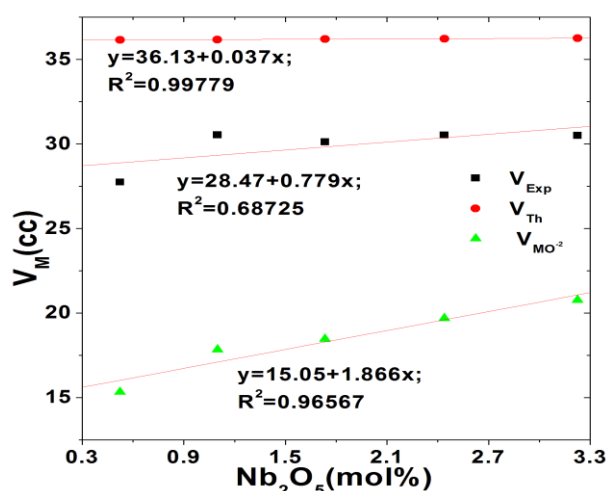


Fig.3 Molar volume and Oxygen molar volume vs. mol% concentration of Niobium doped BGO1:3 glasses. Lines are linear fit to the data. R² values are measure of goodness of fit

On further substitution the rate of increase in these parameters was lowered. This identifies structural saturation of the system on further replacement. **Fig.3** shows that theoretically a rate of 3.7% increase is expected in molar volume with x% increase in niobium however the experimentally observed rate is 77.9% increase while the oxygen molar volume depicted fast increase at the rate of 186.6%. This can be understood as each Niobium ion donates 2.5 oxygen atoms in the network as compared to bismuth and germanium which can collectively provide 2.5 oxygen atoms if in ratio 1:3. Number of oxygen atoms is the main contributing factor. It can be assumed that Nb⁵⁺ is replacing Ge⁴⁺ and modifying the network introducing non-bridging oxygens resulting in extraneous oxygen molar volume.

3.4. Thermal Parameters

3.4.1.DTA- Curves

In **Table2** are included the characteristic transition temperatures, the parameters derived from transformation temperatures and some optical parameters for the system. The glass transition temperature T_g was between $456\text{ }^\circ\text{C}$ - $469\text{ }^\circ\text{C} \pm 2\text{ }^\circ\text{C}$ and onset of crystallization temperature T_x varied $516\text{ }^\circ\text{C}$ - $537\text{ }^\circ\text{C} \pm 2\text{ }^\circ\text{C}$ due to a broad exothermic shoulder. Thermal stability factor $\Delta T = T_x - T_g$ had values $51\text{ }^\circ\text{C} \pm 4\text{ }^\circ\text{C}$ - $69\text{ }^\circ\text{C} \pm 4\text{ }^\circ\text{C}$ showing that except BGO75-Nb10 all other compositions possess comparable stability against devitrification. First crystallization peak apex T_{p1} refers to maximum rate of crystallization for the first crystalline phase that would grow to maturity if temperature is raised above T_g and kept in neighborhood of T_{p1} . Value of T_{p1} shifted from $589\text{ }^\circ\text{C}$ - $624\text{ }^\circ\text{C} \pm 2\text{ }^\circ\text{C}$. It is the main phase that would crystallize if allowed as observed from relative area under the peak. Crystalline peak T_{p2} corresponding to some second crystalline phase was observed in all DTA curves except that for BGO1:3-Nb5 and its temperature decreased non-monotonically from $722\text{ }^\circ\text{C}$ - $694\text{ }^\circ\text{C} \pm 2\text{ }^\circ\text{C}$ with increasing niobium doping. Corresponding fraction crystallized from the composition was less as compared to T_{p1} phase as depicted by relative areas under peaks. Third crystallization peak was observed for all except BGO1:3-Nb10 around $763\text{ }^\circ\text{C} \pm 2\text{ }^\circ\text{C}$ as shown in **Fig.4a**. Another crystalline feature centered about $784\text{ }^\circ\text{C} \pm 2\text{ }^\circ\text{C}$ was observed in 1st two glasses. These features in DTA curves indicate the existence of variety of environments and structural regroupings in the reported glasses. These observations are again witnessing different cation coordination states with varying composition by injecting more niobium. None of the glasses in this study melted in the temperature range scanned. From **Fig.4b** it can be seen that for the pseudobinary glass system reported both T_g and α varied unanimously. The decrease in T_g can be explained as due to modifying character of Nb^{5+} as it caused oxygen molar volume to increase while α decreased as the density increased due to structural rearrangements caused by coordination number changes [9]. Both differentials $T_x - T_g$ and $T_{p1} - T_g$ exhibited similar response to compositional increase in Nb_2O_5 . This may imply that Nb doping pinned nucleation of BGO crystal phase around T_{p1} in BGO75 compositions which correspond to $\text{Bi}_4\text{Ge}_3\text{O}_{12}$ (BGO60) crystallites of stable phase of binary Bi_2O_3 - GeO_2 system [3].

Table2 Characteristic temperatures $T_g, T_x, T_{pi}; i=1-4, \Delta T = T_x - T_g, \Delta T_{p1} = T_{p1} - T_g$, CTE α , refractive index n , direct allowed band gap energy E_g and Urbach energy U_E of niobium doped BGO75 glasses

Nb_2O_5 mol %	T_g $\pm 2^\circ\text{C}$	T_x $\pm 2^\circ\text{C}$	T_{p1} $\pm 2^\circ\text{C}$	T_{p2} $\pm 2^\circ\text{C}$	T_{p3} $\pm 2^\circ\text{C}$	T_{p4} $\pm 2^\circ\text{C}$	$T_x - T_g$ $\pm 4^\circ\text{C}$	$T_{p1} - T_g$ $\pm 4^\circ\text{C}$	$\alpha \times 10^{-6}$ $^\circ\text{C}^{-1}$	n 632nm	$E_{g\text{opt}}$ eV	U_E eV
0.52	464	521	600		764	784	57	136	5.431	1.52	3.244	0.16
1.10	468	537	624	722		785	69	156	5.708	1.52	3.244	0.17
1.73	456	516	589	703	764		60	133	7.333	1.57	2.740	1.39
2.44	469	520	596	664	762		51	127	6.899	1.62	2.701	1.47
3.23	460	518	595	719	763		58	135		1.26	2.630	1.07

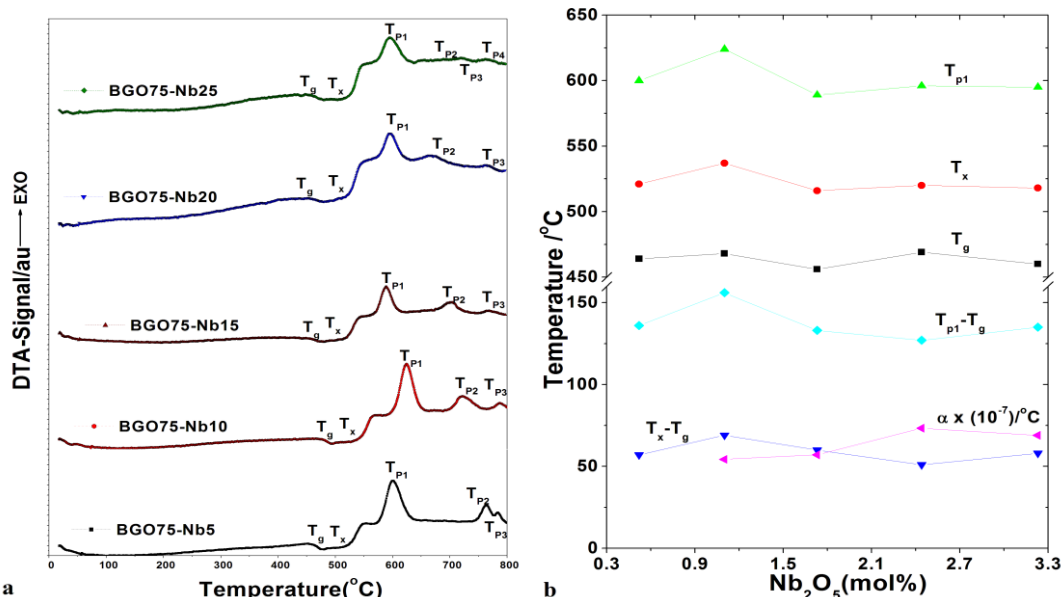


Fig.4a) DTA-curves for Niobium doped BGO75 glasses in temperature range 25-800°C; **b)** Variation of characteristic Temperatures $T_x, T_g, T_{p1}, T_x - T_g, T_{p1} - T_g$ and CTE (α) vs. mol% concentration of Nb_2O_5 for the System

3.4.2. Thermal Coefficient of Linear expansion

The values of thermal coefficient of linear expansion (TCE) α were between $5.431 \pm 0.001 \times 10^{-6} \text{ K}^{-1}$ and $7.333 \pm 0.001 \times 10^{-6} \text{ K}^{-1}$ as reported in the **Table 2** and plotted in **Fig.4b**. TCE is also a structural probe for solid state. An inverse relation was observed between α and density ρ_{exp} , between α and refractive index n while a direct proportion to the molar volume $V_{M exp}$ of the composition. This is expected as an open network expands more on heating. α increased from $5.431 \times 10^{-6} \text{ K}^{-1}$ to $7.333 \times 10^{-6} \text{ K}^{-1}$ as niobium concentration increased from 1.10 to 2.44 mol% the maximum value of α and then decreased to $6.899 \times 10^{-6} \text{ K}^{-1}$ for further increase in niobium concentration to 3.23 mol%. This may be due to co-ordination number change related to structural rearrangement in the network due to larger radius of $\text{Nb}5^+$ cation.

3.5. Optical Parameters

In **Table 2** are included some optical parameters, the refractive index n , direct band gap energy E_g and Urbach energy U_E of the glass system investigated. Refractive index was calculated by reflectance at 632nm using the relation that permittivity ϵ is related to refractive index by

$$\epsilon = n^2 = \frac{1 + \sqrt{R}}{1 - \sqrt{R}} \quad (3)$$

Direct band gap energy E_g and Urbach energy U_E of the glass system were calculated from the UV-NIR absorbance spectra of the glasses using, Lambert-Beer's law that optical absorptivity $\alpha(\nu)$ of the absorbing medium of thickness d traversed is given by

$$\alpha(\nu) = \frac{1}{d} \ln\left(\frac{I_0}{I}\right) \quad (4)$$

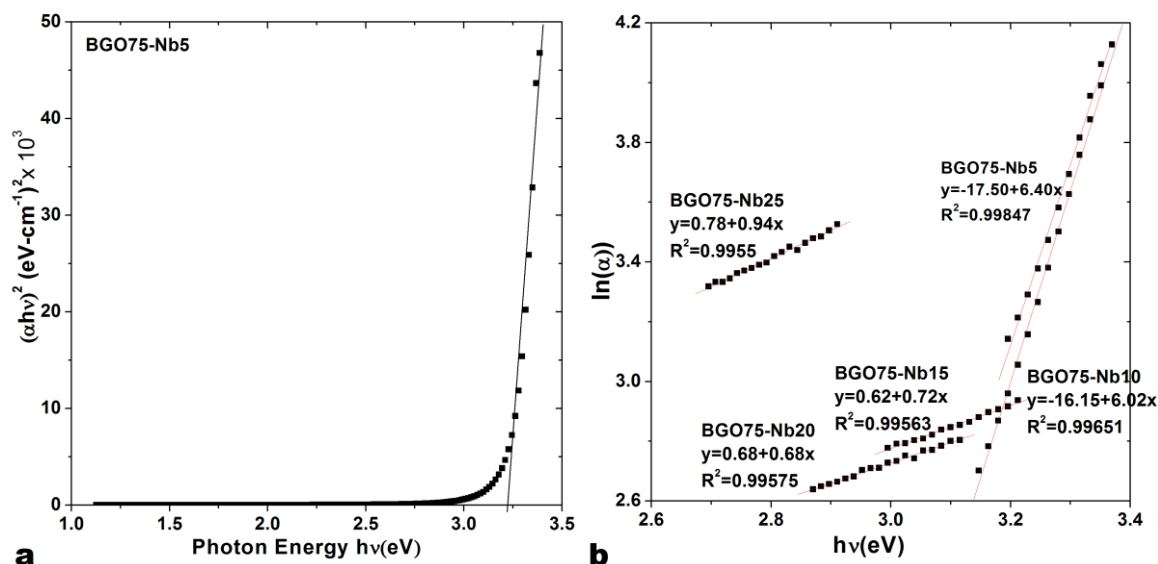


Fig. 5a) $(\alpha h\nu)^2$ vs. $h\nu$ plot for BGO75- Nb5 glass according to Mott and Davis relation for allowed direct band gap energy E_g ; **b)** linear fits to mobility edges according to Urbach Law from Tauc's plot for Niobium doped BGO1:3 glass system

Applying Mott and Davis relation in amorphous solids,

$$\alpha h\nu = B(h\nu - E_g)^r ; r = (2, 3, 1/2, 1/3) \quad (5)$$

Where 2: indirect allowed, 3: indirect forbidden, $1/2$: direct allowed, $1/3$: direct forbidden excitation of electrons across the mobility gap of the amorphous solid. $h\nu$ is the energy of the incident photon of frequency ν . E_g corresponds to bandgap energy and B is band tailing parameter that arises due to amorphous nature of glassy medium in contrast to crystals which have sharp absorption edges.

Fig.5a shows $(\alpha h\nu)^2$ vs. $h\nu$ plots which correspond to $r = 1/2$ and produced a good fit of the data for determination of mobility gap energy E_g , reported in **Table 2** and **Fig.6**. The plot reveals that niobium doped BGO glasses reported are direct bandgap solids and their excitation occurs across allowed states. Optical bandgap energy E_g ranged from 3.244-2.630 eV decreasing with increasing concentration of niobium. This value of band gap is less as compared to simple BGO bandgap energy reported earlier 3.7-3.1 eV by Stentz [12].

Urbach energy U_E of the glasses was calculated from Tauc's Plot (**Fig.5b** $\ln\alpha$ vs. photon energy) as inverse of the slope of the straight line segment near absorption edge. It extended from 0.16-1.47 eV within the band gap. Urbach energy is the measure of the extension of the density of states function within the bandgap of glass from the total order of related crystalline state in terms of band theory of solids. The glass BGO75-Nb20 exhibited the greatest Urbach energy. It was concluded that the system which is 2d ion doped glass obeyed Urbach law as shown in **Fig.5b** similar to glasses reported earlier by P.Petkova for 3d and 4d ions doped BGO glasses [13].

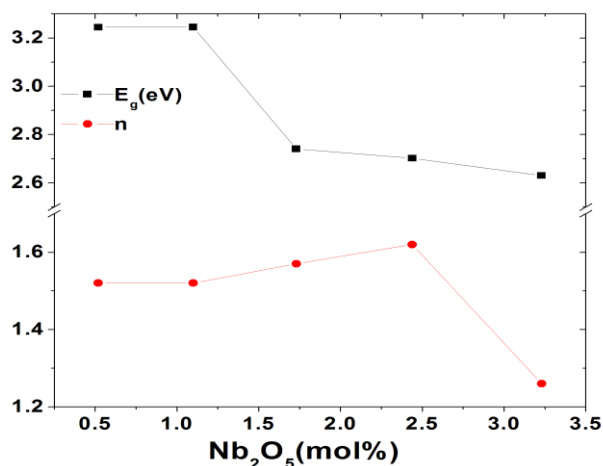


Fig. 6 Variation of Energy band gap E_g and refractive index n

Fig.6 reveals that both optical band gap energy E_g and refractive index n exhibited a plateau against Nb doping initially similar to T_g up to BGO-Nb10. On further doping anion network softened with increase in the molar volume and oxygen molar volume. That may be due to more ionic character of Nb^{5+} ion causing depolymerization and increased NBOs in the network thus assigning modifying behavior to Nb_2O_5 . The increase in NBOs caused T_g to decrease but due to polarizing strength of Nb^{5+} replacing Ge^{4+} ions the refractive index raised continuously up to BGO75-Nb20. On further doping some structural relaxation caused refractive index to decrease. Mobility gap E_g (**Fig.6**) and Urbach edge U_E (**Fig.5b**) also corroborated abrupt change in their values from BGO75-Nb10 to BGO75-Nb15. This may be explained firstly, due to structural co-ordination number change around Bi^{3+} and/or Ge^{4+} cation environments resulting in decreased band gap energy.

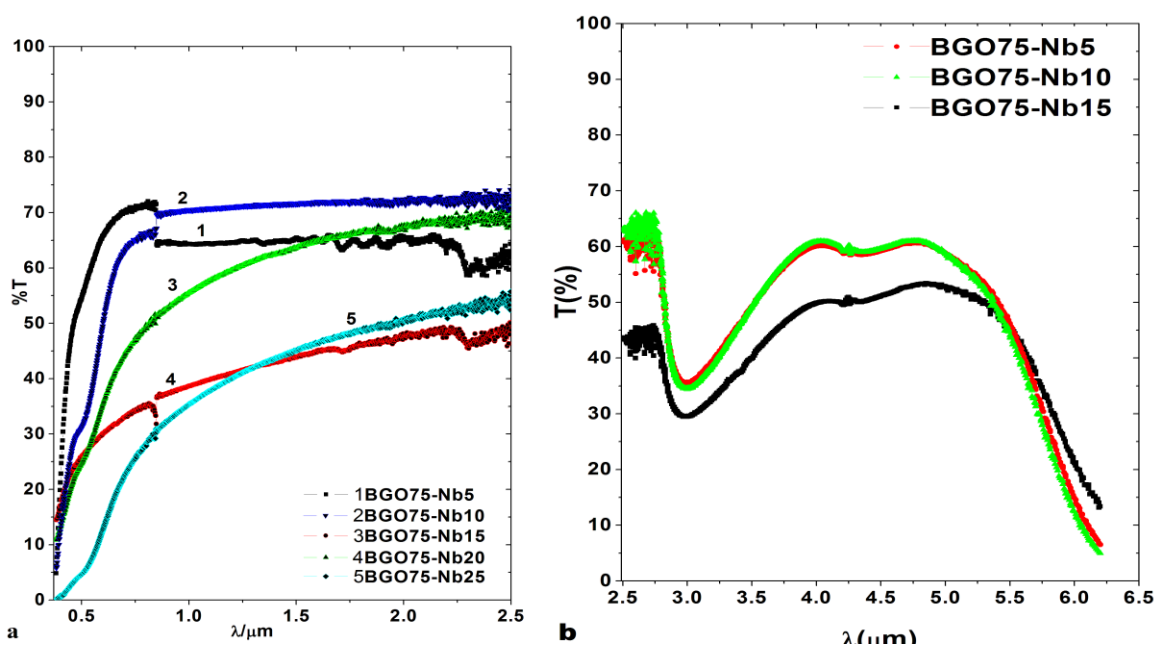


Fig.7a) VIS-NIR transmission spectra of BGO75 glasses; **b)** IR-Transmission spectra of some Niobium doped BGO75 glasses (annealed powder samples) in 2.5-25 μ m range

Secondly, extension of density of states function also tailed into the bandgap due to impurity states of niobium. Extended Urbach exponential edge is manifested due to coupling of phonons and excitation of electrons across these impurity states [8, 11]. The appearance of second and third crystalline features depicted by DTA curves in Fig.2 support this argument.

Fig.7a shows that the optical cut-in wavelength lies at the edge of visible spectrum of electromagnetic radiation for samples of thickness 1-2 mm. %T decreased as thickness of the sample disk increased. Niobium concentration also affected cut in wavelength. It was red shifted as the niobium concentration increased which agrees with the decrease in bandgap energy and increase in Urbach energy that is function of structural arrangement and impurity states [13].

In Fig.7b are reproduced the IR transmission spectra of BGO75 powder samples pre-annealed at 350°C for one hour. Cut off wavelength is extended into MIR region up to 6.4µm. From Fig.7b it is evidenced that IR spectra of BGO glasses are sensitive to compositional variance. The glass BGO75-Nb10 exhibited least absorption and constant percent transmission up to 71 % for disk samples across 0.60 µm -2.5 µm in VIS-NIR region. This transmission efficiency is greater than binary BGO60 glasses [7]. Structural investigation of the lattice vibrations in glass system around cation environment can be performed through IR absorption along with complementary micro Raman spectra in the fingerprint range of 400 cm⁻¹ -1300 cm⁻¹ which is the subject of next investigation. IR transmission in un-annealed powder samples was larger than annealed powders. This may be explained as effect of structural relaxation and more local order in the glasses. The main molecular vibrational absorption in powder samples was corroborated only in longer wavelength region around 1300 cm⁻¹ to 400 cm⁻¹ range.

IV. Conclusions

In ternary glass system BGO75-Nb transparent glasses were successfully molded through normal melt-quench technique. Investigated physical, thermal and optical parameters showed that Niobium initially modifies BGO75 glass system and then imparts structural changes in the glass with increasing mol% concentration of Nb₂O₅. This was revealed by a plateau in its physico-optical and thermal parameters. The value of CTE of glass system was low implying mechanical strength of the niobium doped BGO system. They withstood good thermal and mechanical stability. Optical parameters n , E_g and E_U were affected by both NBOs and polarizing strength of the Niobium cation and therefore varied non-monotonically. The glasses obeyed Urbach Law. The origin of Urbach tails was assigned to impurities/defect states in the bandgap and excitations coupled with lattice phonons. The MIR transmission window was extended to 6.4 µm.

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